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ionic:  $\text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$   
net ionic:  $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$  3.  
 $\text{H}_2\text{SO}_4(\text{aq}) + \text{Mg}(\text{OH})_2(\text{aq}) \rightarrow \text{MgSO}_4(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$  ionic:  $2\text{H}^+(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{Mg}^{2+}(\text{aq}) + 2\text{OH}^-(\text{aq}) \rightarrow \text{Mg}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$  net ionic:  
 $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$  4.  $\text{NH}_4\text{OH}$

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(aq) + HCl (aq)  $\rightarrow$  NH<sub>4</sub>Cl (aq) + H<sub>2</sub>O (l)  
ionic: NH<sub>4</sub><sup>+</sup> (aq) + OH<sup>-</sup>(aq) + H<sup>+</sup> (aq) +  
Cl<sup>-</sup>(aq)  $\rightarrow$  NH<sub>4</sub><sup>+</sup> +

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Author: media.ctsnet.org-Angelika

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strontium acetate and phosphoric acid - 7.  
aluminum sulfate and ammonium sulfide  
8. barium hydroxide and zinc(II)

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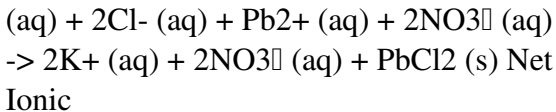
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$2\text{K}^+ (\text{aq}) + 2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+} (\text{aq}) + 2\text{NO}_3^- (\text{aq}) \rightarrow 2\text{K}^+ (\text{aq}) + 2\text{NO}_3^- (\text{aq}) + \text{PbCl}_2 (\text{s})$  Net Ionic Equation:  $2\text{Cl}^- (\text{aq}) + \text{Pb}^{2+} (\text{aq}) \rightarrow \text{PbCl}_2 (\text{s})$  Directions: Write balanced molecular, ionic, and net ionic equations for each of the following reactions. Assume all reactions occur in aqueous solution.

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1a. Identify whether each species  
functions as a Brønsted-Lowry acid or a  
Brønsted-Lowry base in this net ionic  
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reaction that occurs when barium sulfite (s) and excess nitric acid (aq) are combined. Note: Sulfites follow the same solubility trends as sulfates.

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